

## CHAPTER 1

# Chemical Reactions and Equations

- Which of the following is not a physical change?
  - Boiling of water to give water vapour
  - Melting of ice to give water
  - Dissolution of salt in water
  - Combustion of Liquefied Petroleum Gas (LPG)
- The following reaction is an example of a  
$$4\text{NH}_3(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$$
  - displacement reaction
  - combination reaction
  - redox reaction
  - neutralisation reaction
  - (i) and (iv)
  - (ii) and (iii)
  - (i) and (iii)
  - (iii) and (iv)
- Which of the following statements about the given reaction are correct?  
$$3\text{Fe}(\text{s}) + 4\text{H}_2\text{O}(\text{g}) \rightarrow \text{Fe}_3\text{O}_4(\text{s}) + 4\text{H}_2(\text{g})$$
  - Iron metal is getting oxidised
  - Water is getting reduced
  - Water is acting as reducing agent
  - Water is acting as oxidising agent
  - (i), (ii) and (iii)
  - (iii) and (iv)
  - (i), (ii) and (iv)
  - (ii) and (iv)
- Which of the following are exothermic processes?
  - Reaction of water with quick lime
  - Dilution of an acid
  - Evaporation of water
  - Sublimation of camphor (crystals)
  - (i) and (ii)
  - (ii) and (iii)
  - (i) and (iv)
  - (iii) and (iv)

5. Three beakers labelled as A, B and C each containing 25 mL of water were taken. A small amount of NaOH, anhydrous  $\text{CuSO}_4$  and NaCl were added to the beakers A, B and C respectively. It was observed that there was an increase in the temperature of the solutions contained in beakers A and B, whereas in case of beaker C, the temperature of the solution falls. Which one of the following statement(s) is(are) correct?

- (i) In beakers A and B, exothermic process has occurred.
- (ii) In beakers A and B, endothermic process has occurred.
- (iii) In beaker C exothermic process has occurred.
- (iv) In beaker C endothermic process has occurred.

- (a) (i) only
- (b) (ii) only
- (c) (i) and (iv)
- (d) (ii) and (iii)

6. A dilute ferrous sulphate solution was gradually added to the beaker containing acidified permanganate solution. The light purple colour of the solution fades and finally disappears. Which of the following is the correct explanation for the observation?

- (a)  $\text{KMnO}_4$  is an oxidising agent, it oxidises  $\text{FeSO}_4$
- (b)  $\text{FeSO}_4$  acts as an oxidising agent and oxidises  $\text{KMnO}_4$
- (c) The colour disappears due to dilution; no reaction is involved
- (d)  $\text{KMnO}_4$  is an unstable compound and decomposes in presence of  $\text{FeSO}_4$  to a colourless compound.

7. Which among the following is(are) double displacement reaction(s)?

- (i)  $\text{Pb} + \text{CuCl}_2 \rightarrow \text{PbCl}_2 + \text{Cu}$
  - (ii)  $\text{Na}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + 2\text{NaCl}$
  - (iii)  $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$
  - (iv)  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- (a) (i) and (iv)
  - (b) (ii) only
  - (c) (i) and (ii)
  - (d) (iii) and (iv)

8. Which among the following statement(s) is(are) true? Exposure of silver chloride to sunlight for a long duration turns grey due to

- (i) the formation of silver by decomposition of silver chloride
- (ii) sublimation of silver chloride
- (iii) decomposition of chlorine gas from silver chloride
- (iv) oxidation of silver chloride

- (a) (i) only
- (b) (i) and (iii)
- (c) (ii) and (iii)
- (d) (iv) only

9. Solid calcium oxide reacts vigorously with water to form calcium hydroxide accompanied by liberation of heat. This process is called slaking of lime. Calcium hydroxide dissolves in water to form its solution called lime water. Which among the following is (are) true about slaking of lime and the solution formed?
- (i) It is an endothermic reaction
  - (ii) It is an exothermic reaction
  - (iii) The pH of the resulting solution will be more than seven
  - (iv) The pH of the resulting solution will be less than seven
- (a) (i) and (ii)                      (b) (ii) and (iii)  
(c) (i) and (iv)                      (d) (iii) and (iv)
10. Barium chloride on reacting with ammonium sulphate forms barium sulphate and ammonium chloride. Which of the following correctly represents the type of the reaction involved?
- (i) Displacement reaction
  - (ii) Precipitation reaction
  - (iii) Combination reaction
  - (iv) Double displacement reaction
- (a) (i) only                      (b) (ii) only  
(c) (iv) only                      (d) (ii) and (iv)
11. Electrolysis of water is a decomposition reaction. The mole ratio of hydrogen and oxygen gases liberated during electrolysis of water is
- (a) 1:1
  - (b) 2:1
  - (c) 4:1
  - (d) 1:2
12. Which of the following is(are) an endothermic process(es)?
- (i) Dilution of sulphuric acid
  - (ii) Sublimation of dry ice
  - (iii) Condensation of water vapours
  - (iv) Evaporation of water
- (a) (i) and (iii)                      (b) (ii) only  
(c) (iii) only                      (d) (ii) and (iv)
13. In the double displacement reaction between aqueous potassium iodide and aqueous lead nitrate, a yellow precipitate of lead iodide is formed. While performing the activity if lead nitrate is not available, which of the following can be used in place of lead nitrate?
- (a) Lead sulphate (insoluble)
  - (b) Lead acetate
  - (c) Ammonium nitrate
  - (d) Potassium sulphate

14. Which of the following gases can be used for storage of fresh sample of an oil for a long time?

- (a) Carbon dioxide or oxygen
- (b) Nitrogen or oxygen
- (c) Carbon dioxide or helium
- (d) Helium or nitrogen

15. The following reaction is used for the preparation of oxygen gas in the laboratory



Which of the following statement(s) is(are) correct about the reaction?

- (a) It is a decomposition reaction and endothermic in nature
- (b) It is a combination reaction
- (c) It is a decomposition reaction and accompanied by release of heat
- (d) It is a photochemical decomposition reaction and exothermic in nature

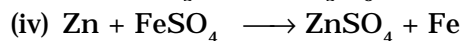
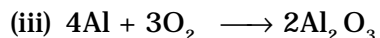
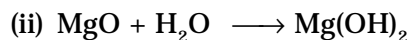
16. Which one of the following processes involve chemical reactions?

- (a) Storing of oxygen gas under pressure in a gas cylinder
- (b) Liquefaction of air
- (c) Keeping petrol in a china dish in the open
- (d) Heating copper wire in presence of air at high temperature

17. In which of the following chemical equations, the abbreviations represent the correct states of the reactants and products involved at reaction temperature?

- (a)  $2\text{H}_2(\text{l}) + \text{O}_2(\text{l}) \rightarrow 2\text{H}_2\text{O}(\text{g})$
- (b)  $2\text{H}_2(\text{g}) + \text{O}_2(\text{l}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
- (c)  $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$
- (d)  $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{g})$

18. Which of the following are combination reactions?



- (a) (i) and (iii)
- (b) (iii) and (iv)
- (c) (ii) and (iv)
- (d) (ii) and (iii)

- 19.** Write the balanced chemical equations for the following reactions and identify the type of reaction in each case.
- Nitrogen gas is treated with hydrogen gas in the presence of a catalyst at 773K to form ammonia gas.
  - Sodium hydroxide solution is treated with acetic acid to form sodium acetate and water.
  - Ethanol is warmed with ethanoic acid to form ethyl acetate in the presence of concentrated  $\text{H}_2\text{SO}_4$ .
  - Ethene is burnt in the presence of oxygen to form carbon dioxide, water and releases heat and light.
- 20.** Write the balanced chemical equations for the following reactions and identify the type of reaction in each case.
- Thermit reaction, iron (III) oxide reacts with aluminium and gives molten iron and aluminium oxide.
  - Magnesium ribbon is burnt in an atmosphere of nitrogen gas to form solid magnesium nitride.
  - Chlorine gas is passed in an aqueous potassium iodide solution to form potassium chloride solution and solid iodine.
  - Ethanol is burnt in air to form carbon dioxide, water and releases heat.
- 21.** Complete the missing components/variables given as x and y in the following reactions
- $\text{Pb}(\text{NO}_3)_2(\text{aq}) + 2\text{KI}(\text{aq}) \longrightarrow \text{PbI}_2(\text{x}) + 2\text{KNO}_3(\text{y})$
  - $\text{Cu}(\text{s}) + 2\text{AgNO}_3(\text{aq}) \longrightarrow \text{Cu}(\text{NO}_3)_2(\text{aq}) + \text{x}(\text{s})$
  - $\text{Zn}(\text{s}) + \text{H}_2\text{SO}_4(\text{aq}) \longrightarrow \text{ZnSO}_4(\text{x}) + \text{H}_2(\text{y})$
  - $\text{CaCO}_3(\text{s}) \xrightarrow{\text{x}} \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$
- 22.** Which among the following changes are exothermic or endothermic in nature?
- Decomposition of ferrous sulphate
  - Dilution of sulphuric acid
  - Dissolution of sodium hydroxide in water
  - Dissolution of ammonium chloride in water
- 23.** Identify the reducing agent in the following reactions
- $4\text{NH}_3 + 5\text{O}_2 \longrightarrow 4\text{NO} + 6\text{H}_2\text{O}$
  - $\text{H}_2\text{O} + \text{F}_2 \longrightarrow \text{HF} + \text{HOF}$
  - $\text{Fe}_2\text{O}_3 + 3\text{CO} \longrightarrow 2\text{Fe} + 3\text{CO}_2$
  - $2\text{H}_2 + \text{O}_2 \longrightarrow 2\text{H}_2\text{O}$

- 24.** Identify the oxidising agent (oxidant) in the following reactions
- $\text{Pb}_3\text{O}_4 + 8\text{HCl} \longrightarrow 3\text{PbCl}_2 + \text{Cl}_2 + 4\text{H}_2\text{O}$
  - $2\text{Mg} + \text{O}_2 \longrightarrow 2\text{MgO}$
  - $\text{CuSO}_4 + \text{Zn} \longrightarrow \text{Cu} + \text{ZnSO}_4$
  - $\text{V}_2\text{O}_5 + 5\text{Ca} \longrightarrow 2\text{V} + 5\text{CaO}$
  - $3\text{Fe} + 4\text{H}_2\text{O} \longrightarrow \text{Fe}_3\text{O}_4 + 4\text{H}_2$
  - $\text{CuO} + \text{H}_2 \longrightarrow \text{Cu} + \text{H}_2\text{O}$
- 25.** Write the balanced chemical equations for the following reactions
- Sodium carbonate on reaction with hydrochloric acid in equal molar concentrations gives sodium chloride and sodium hydrogencarbonate.
  - Sodium hydrogencarbonate on reaction with hydrochloric acid gives sodium chloride, water and liberates carbon dioxide.
  - Copper sulphate on treatment with potassium iodide precipitates cuprous iodide ( $\text{Cu}_2\text{I}_2$ ), liberates iodine gas and also forms potassium sulphate.
- 26.** A solution of potassium chloride when mixed with silver nitrate solution, an insoluble white substance is formed. Write the chemical reaction involved and also mention the type of the chemical reaction?
- 27.** Ferrous sulphate decomposes with the evolution of a gas having a characteristic odour of burning sulphur. Write the chemical reaction involved and identify the type of reaction.
- 28.** Why do fire flies glow at night?
- 29.** Grapes hanging on the plant do not ferment but after being plucked from the plant can be fermented. Under what conditions do these grapes ferment? Is it a chemical or a physical change?
- 30.** Which among the following are physical or chemical changes?
- Evaporation of petrol
  - Burning of Liquefied Petroleum Gas (LPG)
  - Heating of an iron rod to red hot.
  - Curdling of milk
  - Sublimation of solid ammonium chloride
- 31.** During the reaction of some metals with dilute hydrochloric acid, following observations were made.
- Silver metal does not show any change
  - The temperature of the reaction mixture rises when aluminium (Al) is added.
  - The reaction of sodium metal is found to be highly explosive
  - Some bubbles of a gas are seen when lead (Pb) is reacted with the acid.

Explain these observations giving suitable reasons.

32. A substance X, which is an oxide of a group 2 element, is used intensively in the cement industry. This element is present in bones also. On treatment with water it forms a solution which turns red litmus blue. Identify X and also write the chemical reactions involved.
33. Write a balanced chemical equation for each of the following reactions and also classify them.
- Lead acetate solution is treated with dilute hydrochloric acid to form lead chloride and acetic acid solution.
  - A piece of sodium metal is added to absolute ethanol to form sodium ethoxide and hydrogen gas.
  - Iron (III) oxide on heating with carbon monoxide gas reacts to form solid iron and liberates carbon dioxide gas.
  - Hydrogen sulphide gas reacts with oxygen gas to form solid sulphur and liquid water.
34. Why do we store silver chloride in dark coloured bottles?
35. Balance the following chemical equations and identify the type of chemical reaction.
- $\text{Mg(s)} + \text{Cl}_2(\text{g}) \longrightarrow \text{MgCl}_2(\text{s})$
  - $\text{HgO(s)} \xrightarrow{\text{Heat}} \text{Hg(l)} + \text{O}_2(\text{g})$
  - $\text{Na(s)} + \text{S(s)} \xrightarrow{\text{Fuse}} \text{Na}_2\text{S(s)}$
  - $\text{TiCl}_4(\text{l}) + \text{Mg(s)} \longrightarrow \text{Ti(s)} + \text{MgCl}_2(\text{s})$
  - $\text{CaO(s)} + \text{SiO}_2(\text{s}) \longrightarrow \text{CaSiO}_3(\text{s})$
  - $\text{H}_2\text{O}_2(\text{l}) \xrightarrow{\text{U V}} \text{H}_2\text{O(l)} + \text{O}_2(\text{g})$
36. A magnesium ribbon is burnt in oxygen to give a white compound X accompanied by emission of light. If the burning ribbon is now placed in an atmosphere of nitrogen, it continues to burn and forms a compound Y.
- Write the chemical formulae of X and Y.
  - Write a balanced chemical equation, when X is dissolved in water.
37. Zinc liberates hydrogen gas when reacted with dilute hydrochloric acid, whereas copper does not. Explain why?
38. A silver article generally turns black when kept in the open for a few days. The article when rubbed with toothpaste again starts shining.
- Why do silver articles turn black when kept in the open for a few days? Name the phenomenon involved.
  - Name the black substance formed and give its chemical formula.

- 39.** On heating blue coloured powder of copper (II) nitrate in a boiling tube, copper oxide (black), oxygen gas and a brown gas X is formed
- Write a balanced chemical equation of the reaction.
  - Identify the brown gas X evolved.
  - Identify the type of reaction.
  - What could be the pH range of aqueous solution of the gas X?
- 40.** Give the characteristic tests for the following gases
- CO<sub>2</sub>
  - SO<sub>2</sub>
  - O<sub>2</sub>
  - H<sub>2</sub>
- 41.** What happens when a piece of
- zinc metal is added to copper sulphate solution?
  - aluminium metal is added to dilute hydrochloric acid?
  - silver metal is added to copper sulphate solution?
- Also, write the balanced chemical equation if the reaction occurs
- 42.** What happens when zinc granules are treated with dilute solution of H<sub>2</sub>SO<sub>4</sub>, HCl, HNO<sub>3</sub>, NaCl and NaOH, also write the chemical equations if reaction occurs.
- 43.** On adding a drop of barium chloride solution to an aqueous solution of sodium sulphite, white precipitate is obtained.
- Write a balanced chemical equation of the reaction involved
  - What other name can be given to this precipitation reaction?
  - On adding dilute hydrochloric acid to the reaction mixture, white precipitate disappears. Why?
- 44.** You are provided with two containers made up of copper and aluminium. You are also provided with solutions of dilute HCl, dilute HNO<sub>3</sub>, ZnCl<sub>2</sub> and H<sub>2</sub>O. In which of the above containers these solutions can be kept?